1. Give an example of a Combination reaction

2. Give an example of a Decomposition reaction

3. Give an example of a single displacement reaction

4. Give an example of a double displacement reaction

5. Give an example of a Combustion reaction

6. Give an example of an intensive property.

7. When you balance, $\text{Sb}_2\text{S}_3 + \text{HCl} \rightarrow \text{SbCl}_3 + \text{H}_2\text{S}$, the coefficient of HCl is __________

8. Gaining electrons is called __________

9. Identify who got reduced? $\text{Cl}_2 + 2 \text{NaBr} \rightarrow \text{Br}_2 + 2 \text{NaCl}$ __________

10. What is the molar mass of Iron (III) sulfate? __________

11. What is the % by mass of Br in CCl$_2$Br$_2$? __________

12. One mole of Na$_2$SO$_3$ contains how many atoms? __________

13. 0.358 grams of Cu(NO$_3$)$_2$ contains how many moles? __________

14. 1.26 x $10^{22}$ molecules of NaOH corresponds to how many grams of NaOH? __________

15. 2.16 grams of H$_2$SO$_4$ represents how many atoms of oxygen? __________

16. Balance this $\text{C}_6\text{H}_{14} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ name its class __________

17. Balance this $\text{Al} + \text{HCl} \rightarrow \text{AlCl}_3 + \text{H}_2$ Name its class __________

18. Balance this $(\text{NH}_4)_2\text{SO}_4 + \text{Fe(NO}_3)_3 \rightarrow \text{NH}_4\text{NO}_3 + \text{Fe}_2(\text{SO}_4)_3$

19. How many moles of Oxygen are needed to burn 12 moles of C$_6$H$_{14}$ according to the reaction in problem 16? __________

20. 0.478 moles of NH$_4$NO$_3$ is _______ grams and has ________________ molecules.